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Set 3 1. From the
following enthalpy
changes, same $C(s) +$
 $\square O_2(g) \rightarrow \square CO(g) \quad \square H \square$
 $= \square 110.5 \text{ kJ}$ same CO

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$(g) + \frac{1}{2} O_2 (g) \rightarrow CO_2$
 $(g) \quad \Delta H^\circ = -283.0 \text{ kJ}$
calculate the value of
 ΔH° for the reaction C
 $(s) + O_2 (g) \rightarrow CO_2 (g)$.

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30. Compare and

contrast sublimation

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and evaporation. In both processes, the substances become a vapor. During sublimation, the substance goes from the solid phase directly to the vapor phase. During evaporation, particles in a liquid gain enough

States of Matter

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Chemical Bonding 19

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12.3 Arrange the

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following compounds in increasing order of

their boiling points: CH_3CHO , $\text{CH}_3\text{CH}_2\text{OH}$

, CH_3OCH_3 , $\text{CH}_3\text{CH}_2\text{CH}_3$

Ans. The order

is : $\text{CH}_3\text{CH}_2\text{CH}_3 <$

$\text{CH}_3\text{OCH}_3 < \text{CH}_3\text{CHO}$

$< \text{CH}_3\text{CH}_2\text{OH}$

All these compounds

have comparable

molecular masses $\text{CH}_3\text{CH}_2\text{OH}$

undergoes

extensive

intermolecular H-

bonding and thus its

b.pt. is the highest. $\text{CH}_3\text{CH}_2\text{CH}_3$

$\text{CH}_3\text{OCH}_3 < \text{CH}_3\text{CHO}$

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3 CHO is more polar
than CH₃ OCH₃ so ...

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OpenStax 1. The
instantaneous rate is
the rate of a reaction
at any particular point
in time, a period of
time that is so short
that the concentrations

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of reactants and products change by a negligible amount. The initial rate is the instantaneous rate of reaction as it starts (as product just begins to form).

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The notes on Aldehydes and Ketones of class 12 chemistry have been prepared with great care

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keeping in mind the effectiveness of it for the students. This article provides the revision notes of the Aldehydes and Ketones chapter of Class 12 for the students so that they can give a quick glance of the chapter. The chapter has been divided into two articles.

Aldehydes And Ketones Class 12 Notes Chapter 12

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2.10 What role does the molecular interaction play in a solution of alcohol and water? Sol. Alcohol and water both have strong tendency to form intermolecular hydrogen bonding. On mixing the two, a solution is formed as a result of formation of H-bonds between alcohol and H_2O molecules

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but these interactions are weaker and less extensive than those in pure H_2O .

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